



TED (15) - 1004

(REVISION — 2015)

Reg. No.

Signature

FIRST SEMESTER DIPLOMA EXAMINATION IN ENGINEERING/
TECHNOLOGY — APRIL, 2017

ENGINEERING CHEMISTRY - I

(Common to all Diploma Programmes except DCP)

[Time : 3 hours

(Maximum marks : 100)

PART — A

(Maximum marks : 10)

Marks

I Answer the following questions in one or two sentences. Each question carries 2 marks.

1. What are nanomaterials ? Give two examples.
2. What is heterogeneous catalysis ? Write one example.
3. What are conjugate pairs ? Give one example.
4. List any four physical properties of water.
5. What is amalgum ?

(5 × 2 = 10)

PART — B

(Maximum marks : 30)

II Answer *any five* of the following questions. Each question carries 6 marks.

1. Calculate the equivalent weight of (a) NaHCO_3 and (b) Na_2CO_3 .
2. Calculate the molarity of solution prepared by dissolving 2g of NaOH in 500ml of water. What is its normality ?
3. Differentiate between atoms and molecules.
4. What is temporary hardness of water ? How it can be removed ?
5. What are carbon nanotubes ? How they are classified ? Explain.
6. What is bleaching powder ? How it sterilizes water ?
7. What are alloys ? Explain fusion method for the preparation of alloys. (5 × 6 = 30)



PART — C

(Maximum marks : 60)

(Answer *one full* question from each unit. Each full question carries 15 marks.)

UNIT — I

- III (a) Explain any three methods employed for the synthesis of carbon nanotubes. 6
- (b) What are the different properties and applications of carbon nanotubes ? 6
- (c) What is catalysis ? Give two examples. 3

OR

- IV (a) What are the applications of nanomaterials ? 5
- (b) Explain heterogeneous catalysis with two specific examples. 5
- (c) Explain the classification of nanomaterials. 5

UNIT — II

- V (a) Using Lewis concept of acids and bases explain the formation of NH_4^+ and H_3O^+ ions. 5
- (b) Find out the volume of 0.2M HCl required to neutralize 20ml of 0.1M NaOH. 5
- (c) Explain the ionic product of water with necessary equations. What is its value ? 5

OR

- VI (a) Calculate the P^{H} of 0.001M NaOH solution. 6
- (b) What are the buffer solutions ? Explain the types of buffer with examples. 6
- (c) List any three applications of P^{H} ? 3

UNIT — III

- VII (a) Explain the steps involved in the making potable water with the help of flow chart. 6
- (b) Differentiate between hard and soft water. 3
- (c) Explain ion exchange method used for the removal of hardness with equations. 6

OR

- VIII (a) Explain reverse osmosis with necessary diagrams. 6
- (b) Explain the different sterilization methods employed in the production of potable water. 6
- (c) What are the characteristics of potable water ? 3



UNIT — IV

- IX (a) Explain the terms :
(i) Annealing (ii) Quenching (iii) Tempering (iv) Nitriding 6
(b) Explain the various steps involved in powder metallurgy. 6
(c) What are the major purposes of making alloys ? 3
- OR
- X (a) Write down the advantages, disadvantages and applications of powder metallurgy. 8
(b) What are the components of (i) Bronze and (ii) Duralumin ? 4
(c) Write down any three physical properties of metals. 3
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